

# Oxidation States

$\text{Cr}^{3+}$	$\text{CrO}_4^{2-}$	$\text{Cr}_2\text{O}_7^{2-}$	$\text{Cr}_2\text{O}_3$
Cr +3	Cr +6	Cr +3	Cr +3
O -2	O -1	O 0	Cr 0
$\text{ClO}^-$	$\text{ClO}_4^-$	$\text{Cl}_2$	HCl

<b>HCN</b>	<b>O<sub>2</sub></b>	<b>H<sup>+</sup></b>	<b>H<sup>-</sup></b>
<b>O -2</b>	<b>O -2</b>	<b>Cl +1</b>	<b>Cl +7</b>
<b>C -6</b>	<b>N +5</b>	<b>O 0</b>	<b>Cl 0</b>
<b>H<sup>+</sup></b>	<b>Cl -1</b>	<b>N -3</b>	<b>C +2</b>

# Precipitation and Naming Cards

$\text{Li}_2\text{SO}_4$	$\text{LiOH}$	$\text{LiCl}$	$\text{Li}_2\text{S}$
Soluble	Soluble	Soluble	Soluble
$\text{CaSO}_4$	$\text{Ca(OH)}_2$	$\text{Ca(NO}_3)_2$	$\text{Hg}_2\text{Cl}_2$
Not soluble	Not soluble	Soluble	Not soluble

$\text{Hg}_2(\text{NO}_3)_2$	Mercury I nitrate	Mercury II nitrate	Mercury nitrite
Soluble	Mercury nitrate	Lithium sulfide	Lithium sulfate
$\text{Hg}_2\text{SO}_4$	$\text{CS}_2$	Carbon disulfide	$\text{Hg}_2(\text{NO}_3)_2$
Not soluble	$\text{Li}_2\text{S}_2\text{O}_3$	Lithium thiosulfate	$\text{Li}_2\text{S}$